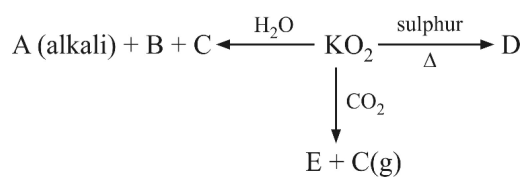


|                                       |                         |                            |
|---------------------------------------|-------------------------|----------------------------|
| Date Planned : __ / __ / __           | Daily Tutorial Sheet-15 | Expected Duration : 30 Min |
| Actual Date of Attempt : __ / __ / __ | Level-3                 | Exact Duration : _____     |

**Paragraph for Q. 159 to 161**



**159.** A is :

- (A)**  $\text{H}_2\text{O}_2$       **(B)**  $\text{O}_2$       **(C)**  $\text{KOH}$       **(D)**  $\text{K}_2\text{CO}_3$

**160.** D is :

- (A)**  $\text{K}_2\text{S}$       **(B)**  $\text{K}_2\text{SO}_4$       **(C)**  $\text{KHSO}_4$       **(D)**  $\text{K}_2\text{S}_2\text{O}_8$

**161.** C is :

- (A)**  $\text{CO}_2$       **(B)**  $\text{CO}$       **(C)**  $\text{O}_2$       **(D)**  $\text{CH}_4$

**162.** Match the chemical properties with the compounds :



**Column I**

- (A)**  $\text{Ca}$   
**(B)**  $\text{CaH}_2$   
**(C)**  $\text{CaO}$   
**(D)**  $\text{CaC}_2$

**Column II**

- (p)** Produces  $\text{H}_2$  on heating with  $\text{H}_2\text{O}$   
**(q)** Produces  $\text{Ca(OH)}_2$  on reaction with  $\text{H}_2\text{O}$   
**(r)** The compound is ionic  
**(s)** Can absorb  $\text{N}_2$  under hot conditions

**163.** An alkali metal nitrate on thermal decomposition gives reddish brown  $\text{NO}_2$  gas. The atomic number of that alkali metal is :



**164.** Number of  $\text{HCO}_3^-$  ions joined by Hydrogen bonding in solid  $\text{KHCO}_3$  is :

